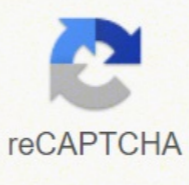




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MO diagram for he2 molecule examples worksheet answers

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MO Electron ^ Configuration: HE2+ LEWIS Structure: Molecular orbital energy diagram Total # of binding electrons # of sigma bonds order Total # of antibonding electrons # of paramagnĂ/diamagnĂ© PI bonds if this molecule exists? When two ^ orbitals are combined, the formation ^ two molecular orbitals occurs. Read Doubts of promising improvement mĂsvedantu Name: Section^ n: Student ID#: Working in groups on these problems. If you get stuck, try asking another group for help. In chemistry, the molecular orbital theory (MO) is a mĂ© all to determine the molecular structure in which electrons are not assigned to individual links between the tomes, but are treated as moving under the influence of the nuclei throughout the molecule. MO Electron ^ Configuration: WITH LEWIS STRUST: Molecular orbital energy diagram Total # of binding electrons # of sigma bonds order Total # of antibonding electrons # of paramagnĂ/diamagnĂ© PI bonds if this molecule exists? Step-by-step response: the theory of molecular orbital assumes that the ^ orbital in molecules lose their identity and the electrons in these molecules are present in new orbital which are commonly called molecular orbital. You should try to answer the questions without referring to your textbook. MO Electron ^ Configuration: O2 LEWIS STRUST: Molecular orbital energy diagram Total # of binding electrons # of sigma bonds order Total # of antibonding electrons # of pi Bonds Paramagnetic/Diamagnetic Should this molecule exist? Molecular orbitals energy diagrams show the relative energies of molecular orbitals. Thus, we can draw the molecular orbital energy diagram as: The stability of molecules can be determined by its bond order which is defined as the half the difference between the number of electrons in bonding molecular orbitals and the number of electrons in antibonding molecular orbitals. H2 Lewis Structure: Molecular Orbital Energy Diagram Total # of bonding electrons # of Sigma Bonds Bond Order Total # of antibonding electrons # of pi Bonds Paramagnetic/Diamagnetic Should this molecule exist? MO electron configuration: Be2 Lewis Structure: Molecular Orbital Energy Diagram Total # of bonding electrons # of Sigma Bonds Bond Order Total # of antibonding electrons # of pi Bonds Paramagnetic/Diamagnetic Should this molecule exist? It is given as:\$Bond\order = \frac{1}{2}[(N_b) - (N_a)]\$And we got from the diagram that Helium has two electrons in its bonding molecular orbital and two electrons in its anti-bonding molecular orbital. Fill in the MO diagram that corresponds to each of the molecules given. Calculate bond order for simple molecules. Practice energy diagrams for molecular orbital theory. MO electron configuration: O22: Lewis Structure: Molecular Orbital Energy Diagram Total # of bonding electrons # of Sigma Bonds Bond Order Total # of antibonding electrons # of pi Bonds Paramagnetic/Diamagnetic Should this molecule exist? MO electron configuration: NO Lewis Structure: Molecular Orbital Energy Diagram Total # of bonding electrons # of Sigma Bonds Bond Order Total # of antibonding electrons # of pi Bonds Paramagnetic/Diamagnetic Should this molecule exist? MO electron configuration: NO+ Lewis Structure: Molecular Orbital Energy Diagram Total # of bonding electrons # of Sigma Bonds Bond Order Total # of antibonding electrons # of pi Bonds Paramagnetic/Diamagnetic Should this "exist? Note: Remember that if the link order value is positive, indicates that the molecule is stable and if the link order value is negative or zero, suggests that the molecule is highly unstable and does not form. MO Electron ^ Configuration ^:

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